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industrially no₂ is an intermediate in the synthesis of nitric acid millions of tons of which are produced each year primarily for the production of fertilizers

acid and base chart lists the strength of acids and bases strongest to weakest in order simple to use laboratory reference chart for scientists researchers and lab technicians

it is acidic in the presence of water it can give origin to nitric and nitrous acid the latter in an environment rich in oxygen will be turned into nitric acid as well

neutral oxides are oxides which are neither acidic nor basic in other words oxides which neither react with acids or with bases are called neutral oxides some of the examples of neutral oxides are a

based on the reactions and properties discussed we conclude that no is an acidic oxide because it can form an acid upon reaction with water while no₂ is classified as a neutral oxide because it does

no nitrogen dioxide is not a basic oxide it is an acidic oxide this is because when nitrogen dioxide reacts with water it forms nitric acid which is an acidic solution

even N_2O nitrous oxide does not display any basic or acidic characteristics it does not form salt when it is reacted with either base or acid so it is also a neutral oxide

NO_2 nitrogen dioxide is classified as a lewis acid a lewis acid is a substance that can accept an electron pair in the case of NO_2 the nitrogen atom has a vacant orbital that can accept an electron

NO_2^- is the nitrite ion a polyatomic ion that consists of one nitrogen atom bonded to two oxygen atoms in the context of acid base theory NO_2^- can be classified as a soft base due to its ability to donate

jun 12 2024 nitrogen dioxide is slightly acidic in nature when dissolved in water it forms nitric acid which can lower the ph of the solution

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